

CHEMICAL FORMULAE & EQUATIONS

Valency	The number of electrons lost, gained or shared by an atom.
Chemical formula	The formula that represents the number & type of the atoms in a molecule.
Atomic groups	A set of atoms joined together that behave like one atom during a chemical reaction and they have their own valency & can't exist solely.
Chemical equations	A set of symbols and chemical formulae, representing the reactants, products and the conditions of the reaction as well.
Reactants	The starting materials in a chemical reaction.
Products	The substances formed from a chemical reaction.
Valence shell	Outermost shell.
Valence electron	Electrons of the outermost shell (group number).
Period number	Number of shells.

- 1) GROUP NUMBERS: I II III IV V VI VII VIII.
- 2) Noble gases have zero valency (no gaining, losing or sharing electrons).
- 3) Prefix: Tells you how many atoms of each element in a formula.
1. Mono- 2. Di- 3. Tri- 4. Tetra- 5. Penta- 6. Hexa- 7. Hepta- 8. Octa- 9. Nona- 10. Deca-.
- 4) Non-metal ions take the suffix (ide).
Chloride / Nitride / Oxide.
- 5) Methane → CH₄.
- 6) Zinc has a valency of 2.

ATOMIC GROUPS

ATOMIC GROUP	SYMBOL	CHARGE	VALENCY
Hydroxide	OH	-ve	1
Hydrogen carbonate	HCO ₃	-ve	1
Nitrite	NO ₂	-ve	1
Nitrate	NO ₃	-ve	1
Ammonium	NH ₄	+ve	2
Carbonate	CO ₃	-ve	2
Sulfate	SO ₄	-ve	2
Sulfite	SO ₃	-ve	2
Phosphate	PO ₄	-ve	3
Phosphite	PO ₃	-ve	3

MOLES

Moles	A unit of measurement used in chemistry to express the amounts of a chemical substance.
Moles	The quantity of a substance that contains 6×10^{23} .
Moles	The number of particles which is equal to the number of atoms in 12g Carbon.
Moles	The mass in grams which contains the Avogadro's constant number of particles.
Avogadro's constant	The number of particles in one mole of a substance.
Relative atomic mass	The average mass of the isotopes of elements.
Relative molecular mass	The sum of atomic masses of the elements in a compound.



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