

Polyatomic Ions

Ion	Name	Charge
NH ₄	Ammonium	1+
Hg ₂	Mercury(1)	2+
NO ₃	Nitrate	1-
NO ₂	Nitrite	1-
HSO ₄	Bisulfate	1-
OH	Hydroxide	1-
CN	Cyanide	1-
H ₂ PO ₄	Dihydrogen Phosphate	1-
NCS	Thiocyanate	1-
HCO ₃	Bicarbonate	1-
ClO	Hypochlorite	1-
ClO ₂	Chlorite	1-
ClO ₃	Chlorate	1-
ClO ₄	Perchlorate	1-
C ₂ H ₃ O ₂	Acetate	1-
MnO ₄	Permanganate	1-
SO ₃	Sulfite	2-
SO ₄	Sulfate	2-
HPO ₄	Hydrogen Phosphate	2-
CO ₃	Carbonate	2-
Cr ₂ O ₇	Dichromate	2-
CrO ₄	Chromate	2-
O ₂	Peroxide	2-
C ₂ O ₄	Oxalate	2-
TeO ₂	Hypotellurite	2-
TeO ₃	Tellurite	2-
TeO ₄	Tellurate	2-
TeO ₅	Pertellurate	2-
PO ₄	Phosphate	3-
PO ₃	Phosphite	3-

Strong Acids

Name	Formula
Hydrochloric Acid	HCl
Hydrobromic Acid	HBr
Hydroiodic Acid	HI
Chloric Acid	HClO ₃
Perchloric Acid	HClO ₄
Nitric Acid	HNO ₃
Sulfuric Acid	H ₂ SO ₄

Strong Bases

Name	Formula
Lithium Hydroxide	LiOH
Sodium Hydroxide	NaOH
Potassium Hydroxide	KOH
Rubidium Hydroxide	RbOH
Calcium Hydroxides	Ca(OH) ₂
Strontium Hydroxide	Sr(OH) ₂

Other Acids

Name	Formula
Acetic Acid	CH ₃ COOH

Water Energies

Delta H Sublimation	46.6 kJ/mol
Delta H Fusion	6.02 kJ/mol
Delta H Vaporization	40.6 kJ/mol

Intermolecular Forces

Dipole - Dipole Forces

Neutral Polar molecules attract each other when the positive end of one molecule is near the negative end of another. Smaller molecules have a higher dipole-dipole attractive forces.

London Dispersion Forces

Intermolecular Forces (cont)

Neutral non-polar molecules due to the instantaneous distribution of electrons.

Temporary dipole on one atom induces a similar dipole on adjacent atom causing the atoms to be attracted to each other.

Hydrogen Bonds

An attraction between the hydrogen atom in a polar bond that is bonded to an electronegative atom and the lone pairs of electrons on another atom. Stronger than dipole-dipole or London Dispersion forces.

Molarity

Molarity (M) = (moles of solute)/(Liters of Solution)

solubility rules



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Not published yet.

Last updated 29th September, 2021.

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