

Changing state

The temperature will remain constant during the change of state.

The melting and freezing point are the same temperature.

Mixtures

Homogeneous mixtures

Heterogeneous mixtures

Group 1

Alkali

Silvery-white coloured

Metals

Relatively low boiling points

They react with water to produce hydrogen gas and heat.

The heat produced can ignite (burn) the hydrogen gas produced.

As you move down the group, the reactions become more violent.

Group 2

Alkaline

Relatively high boiling points

The reactions are less violent than group 1 elements.

Group 17

Halogens

diatomic

very reactive with metals

F₂ and Cl₂ pale green gases

Br₂ red brown liquid

I₂ grey solid

Group 18

nobel gases

unreactive

Covalent Compounds

non-metal atoms chemically bonded together

Ionic Compounds

monatomic ion

polyatomic ion

single atom that has a charge

group of atoms with a charge

Predicting bond type

Metals and non-metals ionic bonds

Non-metals covalent bonds

Reactions of Metals

Reactions of Metals

Reactions of

Acids and Bases

Metal + Oxygen Metal

Acid + Carbonate

oxide Metal + Water

Metal salt +

Metal hydroxide +

Water + Carbon

Hydrogen Metal + Acid

dioxide Acid +

Metal salt + Hydrogen

Base Metal salt +

Gas

Water

Metal + Oxygen Metal oxide

Metal + Water Metal hydroxide + Hydrogen

Metal + Acid Metal salt + Hydrogen Gas

Corrosion

name 3 methods to reduce corrosion

Painting Tin coating Chrome plating

Enamelling Plastic coating Galvanizing

(Zinc coating) Anodizing Alloying



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