Cheatography

Chem Rules and Formulas Cheat Sheet

by Schutte via cheatography.com/30544/cs/9121/

Chapter 1 Formulas

T(°C) = T(K) - 273.15
T(°F) = 9/5 [T(°C)] + 32
Density = mass/volume
average atomic weight = (exact mass of isotope) (fractional abundance of isotope) + ...

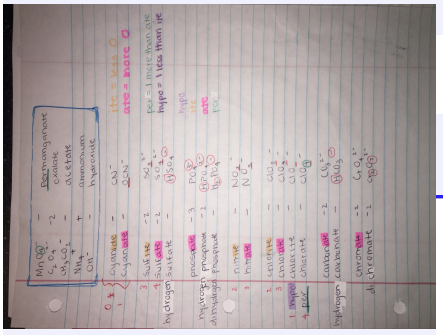
Chapter 1 Conversions

1 in	2.54 cm
1 mi	5280 ft

Diatomic Molecules

H2
N2
O2
F2
Cl2
Br2
I2

Polyatomic Ions



Chapter 4 Oxidation Numbers

element by itself	0
Group 1A	+1
Group 2A	+2
Halogen	usually -1 / positive with Oxygen
Monatomic Ion	ion charge
H	+1 with nonmetals / -1 with metals
O	usually -2 / -1 in peroxide (O2)
F	ALWAYS -1

Chapter 4 Solubility Rules

exceptions	soluble ionic compounds	insoluble ionic compounds
	Na^+	OH^-
	K^+	O^{2-}
	$NH4^+$	S^{2-}
	$NO3^-$	$CO3^{2-}$

Chapter 4 Solubility Rules (cont)

exceptions	soluble ionic compounds	insoluble ionic compounds
	$CH3CO2^-$	$PO4^{3-}$
	$ClO3^-$	
	$ClO4^-$	
Pb^{2+}, Ag^+, $Hg2^{2+}$	Cl^-	
Pb^{2+}, Ag^+, $Hg2^{2+}$	Br^-	
Pb^{2+}, Ag^+, $Hg2^{2+}$	I^-	
Ca^+, Sr^+, Ba^{2+}, Pb^+	F^-	exceptions
Ca^+, Sr^+, Ba^{2+}, Pb^+	$SO4^{2-}$	Na^+, K^+, Ba^{2+}
		Na^+, K^+, Ba^{2+}
		Na^+, K^+, $NH4^+$, Ba^+
		Na^+, K^+, $NH4^+$

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